

AQA, OCR, Edexcel

GCSE Science

GCSE Chemistry

**Electrolysis
Questions**

M M E

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Total Marks: /33

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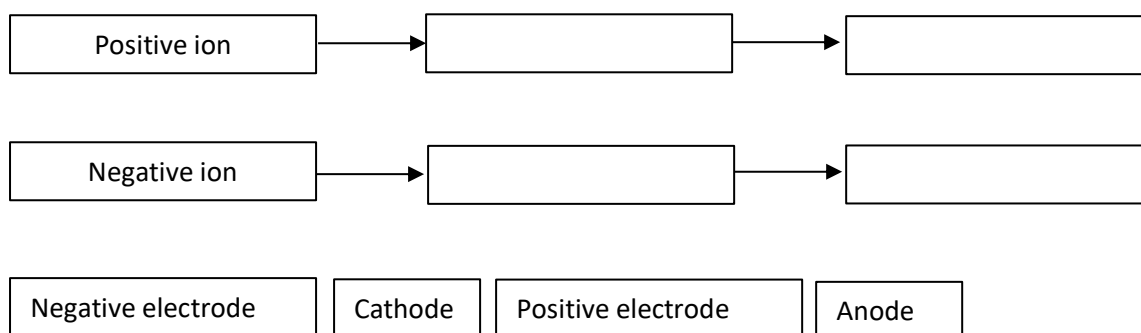
Q1: When an ionic compound is melted or dissolved in water, what happens to the ions?

(1 mark)

Q2: Define an electrolyte.

(2 marks)

Q3: Passing an electric current through electrolytes causes the ions to move to electrodes. Match up the boxes.



(4 marks)

Q4: What is the process called?

(1 mark)

Q5: If an ionic compound (a non-metal and a metal) is electrolysed at which electrode will the metal be produced and the non-metal.

Metal _____

Non-metal _____

(2 marks)

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Q6: When is electrolysis used to extract metals?

(2 marks)

Q7: What is the molten mixture used to manufacture aluminium through electrolysis?

(2 marks)

Q8: What metal is used at the positive electrode?

(1 mark)

Q9: Why is a mixture used as the electrolyte?

(2 marks)

Q10: Why must the positive electrode be constantly replaced?

(2 marks)

Q11: When an aqueous solution is electrolysed using inert electrodes, the ions that are discharged are dependent upon what?

(2 marks)

Q12: If the metal is more reactive than hydrogen, what is produced at the negative electrode?

(1 mark)

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Q13: At the positive electrode, what is produced?

(1 mark)

Q14: Why does this happen in aqueous solution?

(2 marks)

Q15: Complete the following sentences.

During electrolysis, at the _____ (negative electrode), _____ charged ions _____ electrons.

These reactions are _____ reactions.

At the _____ (positive electrode), _____ charged ions _____ electrons.

These reactions are _____ reactions.

cathode	negatively	gain	oxidation
anode	positively	lose	reduction

(8 marks)